# Mark Scheme - AS 2.3 The Wider Impact of Chemistry

1 (a) Benefits:
Stops fossil fuels from running out
Reduces $CO_2$ emissions / greenhouse emissions / global warming / effect of global warming
Reduces $SO_2$ emissions / acid rain
There will be an investment in new technology

Difficulties:
Dependence on fossil fuel/Unlikely to meet current demand
Renewable energy currently more expensive
Reliability of supply from renewables
Major development in energy efficiency technologies required
Opposition by vested interests
(Maximum 3 marks from list, but need examples of both) (3)

Consideration and discussion of benefits/difficulties (1) [4]

*QWC* Legibility of text; accuracy of spelling, punctuation and grammar, clarity of meaning [1]

(b) (i) I As temperature increases yield decreases
As pressure increases yield decreases [1]

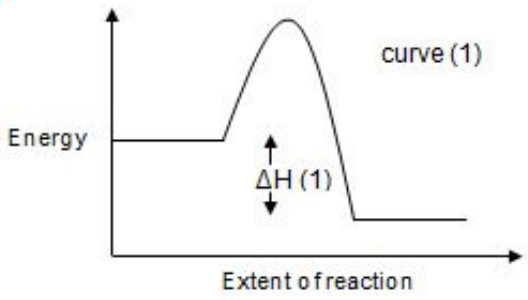
II As temperature is increased, equilibrium moves to the left (1)
Therefore forward reaction is exothermic (1)
As pressure is increased, equilibrium moves to the left (1)
Therefore more gas moles in products (1) [4]
*QWC* The information is organised clearly and coherently, using specialist vocabulary where appropriate [1]

(ii) If temperature is too low, then reaction is too slow (1)
If temperature is too high, yield is too low (1)
Compromise temperature – acceptable rate and yield (1)
(Accept any two points) [2]

(iii) Heterogenous catalyst [1]

(iv) Lower temperatures could be used (1)
Less energy consumption / increased yield (1)
Equilibrium could be reached more quickly (1)
(Accept any two points) [2]

**(v)**




[2]

(vi) $\Delta H = E_f - E_b$ [1]

**Total [19]**

2 (a) (i) Temperature: 298K / 25°C (1) Pressure: 1 atm / 101.325 kPa or 100 kPa (1) [2]

(ii) Hydrogen gas is an element in its standard state [1]

(iii) $\Delta H = \Delta H_f (C_5H_{12}) + 5\ \Delta H_f (H_2O) - 5\ \Delta H_f (CO) - 11\Delta H_f (H_2)$ (1)

$\Delta H_f (C_5H_{12}) = -1049 - 5\ (-286) + 5\ (-111)$ (1)

$\Delta H_f (C_5H_{12}) = -174\ kJ\ mol^{-1}$ (1) [3]

(b) (i) Catalyst in different state to reactants [1]

(ii) Catalysts provide an alternative route (1) with a lower activation energy (1) [2]

(iii) Lower temperature or less time so less energy needed / Can make alternative production method possible with sustainable starting materials or less waste products [1]

(iv) At higher temperatures particles have more energy (1)

More collisions have energy above activation energy (1)

*(Can obtain these two marks from correctly labelled Boltzmann energy distribution plot with two temperature lines (1) and Activation energy (1))*

Successful collisions occur more frequently (1) – 3 max [3]

*QWC: selection of a form and style of writing appropriate to purpose and to complexity of subject matter* [1]

(c) (i) No effect (1)

Same number of (gas) molecules on both sides of reaction (1) [2]

(ii) Lower yield of hydrogen (1)

Reaction shifts in endothermic direction to (try to counteract increase in temperature) (1) [2]

(iii) No effect [1]

**Total [19]**

34 (a) killing marine life / killing trees [1]

(b) (i) either gas syringe or inverted burette attached to sealed vessel [1]

(ii) different surface area would affect rate of reaction [1]

(iii) concentration / volume / nature of acid (1)

temperature (1) [2]

(c) (i) increasing pressure will shift the reaction to side with fewer gas molecules (1)

increasing yield of $SO_3$ (1) – reason must be given [2]

(ii) I increasing temperature shifts equilibrium in endothermic direction (1)

as $SO_3$ yield is decreased forward reaction must be exothermic (1) [2]

II increasing temperature increases energy of particles (1)

more collisions have energy above activation energy (1)

successful collisions occur more frequently (1)

*can gain first two points from labelled Boltzmann distribution curve* [3]

III e.g. iron in production of ammonia or any valid example [1]

(d) (i) atom economy = 100% [1]

(ii) any two points from:

lower pressure used in B (1)

methanol is a renewable starting material (1)

higher atom economy in B or less waste in B (1)

[ignore reference to cost]
2 max [2]

(iii) no effect on position of equilibrium [1]

Total [17]

4 (a) Oil is non-renewable / will run out (1)

Contribution of $CO_2$ to global warming (1)

Oil has other important uses (1) [2]

(Maximum 2 marks)

(b) (i) Power stations / fossil fuels used to generate the electricity needed to make $H_2$ (1)

Resulting in $CO_2$ formation (global warming) / acid rain (1)

Manufacture of car produces pollution (1) [2]

(Maximum 2 marks)

*QWC* Legibility of text; accuracy of spelling, punctuation and grammar, clarity of meaning [1]

(ii) Disagree, no fuel is 100% safe /
petrol can burn explosively
(Accept agree if valid reason given e.g. in terms of lives being lost) [1]

(c) (i) Hydrogen since frequency is inversely proportional to wavelength / smaller wavelength [1]

(ii) Hydrogen since energy is proportional to frequency / greater frequency / E = hf [1]

(d) In Ne greater shielding of *outer* electron (1) outweighs larger nuclear charge (1) / He has greater effective nuclear charge (1) / He *outer* electron closer to nucleus (1)
- max 1 if no reference to *outer* electron [2]

(Maximum 2 marks)

(e) (i) $^{218}Po$ [1]

(ii) Since radon is a gas / inhaled, α particles will be given off in the lungs (which may cause cancer) [1]

Total [12]

5 (a) (i) They are both elements in their standard states. [1]

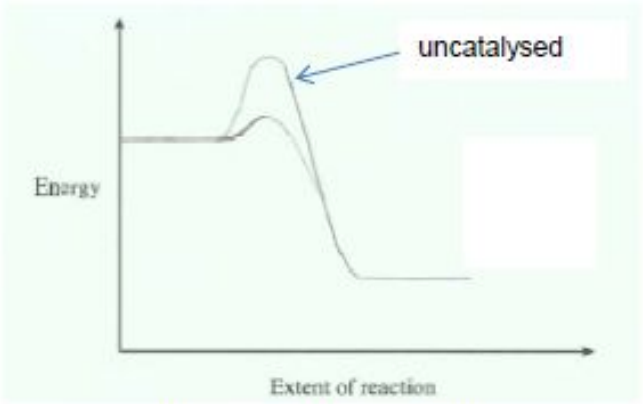
(ii) $\Delta H = \Sigma \Delta H_f$ products $- \Sigma \Delta H_f$ reactants (1)

$= (-286 + 0) - (-368 + 0)$

$= -286 + 368 = (+)82$ (kJ mol$^{-1}$) (1) [2]

or by a cycle where correct cycle drawn (1) correct answer (1)

(b) (i)




exothermic profile drawn (1)
uncatalysed / catalysed line labelled (1) [2]

(ii) I number of moles of benzene = 2000 [1]

II mole ratio is 1 : 1 (1)

$\therefore$ moles of phenol produced $= \frac{2000 \times 95}{100} = 1900$ (1)

mass $= M_r \times$ number of moles $= 94 \times 1900 = 178.6 / 179$ kg (1)

*alternatively*

78 (g / kg) of benzene gives 94 (g / kg) of phenol (1)

$\therefore$ 1 (g / kg) of benzene gives 94/78 (g / kg) of phenol
$\therefore$ 156 (kg) of benzene gives $94 \times 156/78$ (kg) of phenol = 188 (kg) (1)
but 95% yield $\therefore \frac{188 \times 95}{100} = 178.6 / 179$ (kg) (1) [3]

(iii) Look for at least four relevant positive points [4] e.g.

- the process uses a (heterogeneous) catalyst, which can easily be separated from the gaseous products (thus saving energy)
- the only other product of the reaction is gaseous nitrogen, which is non-toxic / safe / not a harmful product
- the process uses nitrogen(I) oxide which is used up, rather than being released into the atmosphere from the other process (and causing global warming)
- the process is exothermic and the heat produced can be used elsewhere
- a relatively moderate operating temperature reduces overall costs
- high atom economy

Legibility of text; accuracy of spelling, punctuation and grammar;
clarity of meaning QWC [1]
Total [14]

6 (a) $K \rightarrow 1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ (1)
There is one outer electron and the loss of this electron gives a stable potassium ion with a full outer shell/ ion more stable than the atom (1) [2]

(b) (i) $\Delta T = 4.8\ ^{o}C$ (1)

$$\Delta H = -\frac{250 \times 4.2 \times 4.8}{0.125} = -40320\ J\ mol^{-1} / -40.3\ kJ\ mol^{-1}$$ (2) [3]

✓ for negative sign
✓ correct value with relevant units

(ii) e.g. The volume used was not precise in measurement as the readings on a beaker are only approximate (1)
The experiment was performed in a beaker and this was not insulated and heat was lost to the surroundings (1) [2]

there may be other acceptable answers here, for example based on slow dissolving

(c) (i) 0.050 [1]

(ii) $(0.050 \times 24.0) = 1.20\ (dm^3)$ [1]

(iii) $\%\ v/v = \frac{1.20 \times 0.001 \times 100}{2}$ (1) $= 0.06$ (1) [2]

(d) An increase in the concentration of (aqueous) carbon dioxide causes the position of equilibrium to move to the right. (1)
This causes calcium carbonate to become aqueous calcium (and hydrogencarbonate) ions / dissolve (1)
weakening shells / causing difficulty in formation of shells (1) [3]

*Organisation of information clearly and coherently; using specialist vocabulary where appropriate* QWC [1]

**Total [15]**

7 (a) Enthalpy change when one mole of a compound is formed from its (constituent) elements (1)
in their standard states / under standard conditions (1) [2]

(b) (i) $H_2 + \frac{1}{2}O_2 \rightarrow H_2O$ [1]

(ii) $-242 = 436 + 248 - 2(O—H)$ (1)
$2(O—H) = 926$
$O—H = 463\ kJ\ mol^{-1}$ (1) [2]

(c) (i) I. Burning hydrogen will not produce $CO_2$ (or $SO_2$) as pollutants [1]

II. Hydrogen is very flammable, storing as $MgH_2$ is safer / $MgH_2$ is solid therefore volume occupied by given amount of hydrogen is less [1]

(ii) If the $MgH_2$ is not kept dry, hydrogen will be formed and there could be a potential explosion [1]

(iii) Moles $MgH_2 = \frac{70000}{26.32} = 2659.6$ (2660) (1)

Moles $H_2 = 5319.2$ (5320) (1)

Volume $H_2 = 1.28 \times 10^5\ dm^3$ (1) [3]

(d) (i) An increase in temperature would decrease the yield and an increase in pressure would increase the yield [1]

(ii) Forward reaction is exothermic so equilibrium shifts to the left as temperature is increased (1)

More gaseous moles on the l.h.s. so equilibrium shifts to the right as pressure is increased (1) [2]

(e) Lower temperatures can be used (1)
Energy costs saved (1)
More product can be made in a given time (so more can be sold) (1)
Enable reactions to take place that would be impossible otherwise (1)
Less fossil fuels burned to provide energy (so less $CO_2$ formed) (1)
(any 3 of above) [3]

*QWC Legibility of text; accuracy of spelling, punctuation and grammar, clarity of meaning* [1]

**Total [18]**